

# Operating Instructions & Technical Specifications

## **directION Copper Combination ISE.**

The *directION* Copper combination ion selective electrode has a solid state crystalline membrane with an integral driTEK reference. The electrode is designed for the detection and analysis of Copper ions in aqueous solutions and is suitable for use in the field ,the laboratory and in on line analyzers.

### **Installation**

Connect the ISE to the mV or ion meter.

Remove the black protective cap and keep it in a safe place.

The ISE can be used immediately but pre soaking for 10 minutes in a 100 ppm Copper solution is recommended.

The ionic strength of the standards and solutions should be kept constant between all standards and samples. This is achieved by the simple addition of an Ionic strength adjustment buffer (ISAB). Potassium Nitrate is ideal. A typical addition would be 2 ml of 1 molar ISAB to 100 ml of standard and sample.

For low level measurements below around 50 ppm in relatively pure samples no ISAB is needed.

No temperature correction is necessary however standards and samples should be measured at the same temperature.

Begin calibration from the lowest concentration standard to avoid cross contamination. Calibration should cover the anticipated range of the samples. Standards should be of similar matrix to the samples being measured. If this is not possible and the results show an offset i.e. are consistently high or low then you should consider using an incremental technique. These techniques are outlined in this manual.

Rinse tip with deionised water between measurements.

Avoid strongly acidic or alkaline samples, and organic solvents.

### **Storage and Maintenance**

After use rinse with deionised water , replace protective cap and store dry in its box.

If performance becomes sluggish carefully polish the membrane with a fine emery cloth or polishing strip, rinse with deionised water and immerse the tip in a 100 ppm Copper solution for 10 minutes.

Avoid prolonged exposure to solutions of high concentration.

The Membrane surface should appear shiny and greyish blue in appearance.

## Specification;

Copper ( Cu 2+) is a divalent cation. The electrode is not suitable for measuring monovalent Copper ions.

Refer to the table below for the full specification:

### Catalogue number 3227      Copper *directION* Combination ISE

Parameter	Specification
Overall length	155 mm
Body Diameter	12 mm
Cap Diameter	16mm
Connector	BNC
Cable length	1000 mm
Resistance at 25 Deg C	< 2.5 Meg Ohm
Concentration Range	0.006 to 64,000 ppm
Slope	23 to 29 mV per decade
Potential Drift	2 mV per day
Operating pH range	2 to 7
Temperature range	5 to 50 Deg C
Endpoint time	Typically 30 to 60 seconds
ISAB	2 Molar KNO3
Interferences: Ions with coefficients above 0.001.	Mercury Silver Sulphide

## Recommendations for successful analysis.

*directION* combination ISE's can be used with any pH/mV meter or Ion meter. If the meter does not have a BNC socket and you have a BNC electrode please contact your *directION* distributor who will arrange to have the correct plug fitted. Adapters are also available if the same electrode has to be used on more than one meter.

**Meters** with a 0.1 mV resolution are recommended whilst dedicated Ion meters will provide direct concentration readouts saving time and effort in constructing calibration curves and performing calculations. Your *directION* distributor can advise on the most suitable meter.

**Magnetic stirrer/stirrer bars** are recommended for laboratory analysis. Please operate at the lowest constant speed available.

**Semi-logarithmic (4-cycle) graph paper** is required for preparing calibration curves when you are using a mV meter.

## Required Solutions

**Distilled or deionised water** will be required to prepare Standards, ISABs and to rinse the electrode between measurements.

**1000 ppm Stock Standard solution.** Used for preparation of Standards. (Prepared by customer)

**ISAB.** Used to adjust the Ionic strength of all standards and samples. Typical addition is 1 ml of 4 Molar ISAB to 50ml of all standards and samples.

## Operation

1. Connect the electrode to the meter being used for analysis
2. Prepare a series (at least 2) of standards that bracket the expected sample concentration. This is best done by serial dilution of the stock solution. Ideally standards should be a decade in concentration apart e.g. 1, 10, and 100 ppm.
3. Dispense 50 ml of each standard into analytically clean beakers (100 to 150 ml size is perfect)
4. Add ISAB/TISAB in the appropriate ratio. As a guide with sample concentrations in the 1 to 1000 ppm range 1ml of a 4 Molar ISAB to 50 ml sample is satisfactory. For TISAB(Fluoride analysis only) please read the label.
5. Rinse the electrode with deionised water and blot dry with a lint free cloth and place in the lowest standard. When the reading is stable record the mV value.
6. Repeat step 6 for all subsequent standards proceeding from lowest to highest.
7. Plot a calibration curve on semi log paper using mV values on the linear Axis and concentration on the log scale.
8. Rinse the electrode in deionised water and blot dry. Place the electrode in the sample and record the stable mV value.
9. Using the calibration curve determine the unknown sample concentration.

## Hints and tips.

1. Ensure that the temperature of all standards and samples are the same to reduce errors.
2. Using a magnetic stirrer for laboratory analysis is recommended but not essential. It is however important to have the stirrer set on a low constant speed which must be reproducible for all measurements.
3. Prior to sample measurement ensure that the electrode is thoroughly rinsed with deionised water. It is worth performing this rinse twice given the possibility of carryover being greatest in high concentration solutions.
4. Prepare standards by serial dilution.
5. Make sure your electrode is conditioned by leaving the tip in the lowest concentration standard for 1 hour prior to analysis.

## Methods of Analysis.

We have described direct potentiometry above. This method is simplified by using a direct reading ion meter. There are several other methods, which are useful. For full advice or copies of applications method sheets please contact [info@edt.co.uk](mailto:info@edt.co.uk)

**Known Addition:** An incremental technique where the potential of the sample solution is measured followed by addition of a small volume of a higher concentration standard solution. The new potential is measured and from difference in the two values, and using the known electrode slope, the unknown concentration is determined,

This method is ideal for samples whose matrix is not entirely clean or aqueous. In these instances comparisons with clean standards is not appropriate thereby making direct potentiometry unsuitable. Known addition works because both standard and sample are measured in the same matrix.

Typical sample volume is 50 ml, typical standard volume is 5 ml. The standard should be approximately 100 times the sample concentration for accurate analysis.

**Sample Addition:** An incremental technique where the potential of a dilute standard solution is measured followed by the addition of a small volume of more concentrated sample. The new potential is recorded and the difference noted. Using this value (and the predetermined electrode slope) the unknown concentration is determined.

This method is ideal for dirty or viscous samples with an awkward matrix. The sample however needs to be relatively concentrated i.e. at least 100 times the Electrodes linear detection limit. The analysis does have the benefit of only requiring a small volume.

The sample matrix is basically broken down by dilution with the standard and therefore analysis is carried out in the same media.

**End Point Titration:** Flow Plus combination ISE's are ideal end point indicators and will produce a significant potential change at the equivalence point. The Ion in question must be contained in the titrand or the titrant and must therefore be in excess or absence at the end point.

### **Applications Advice:**

There are three types of direction ISE all of which have advantages in certain applications or circumstances. It is important to select the correct ISE for your application. The typical situations are listed below with the most appropriate ISE type. When in doubt contact your local direction distributor or e mail [info@edt.co.uk](mailto:info@edt.co.uk) for further advice or support.

#### **Half Cell ISE with separate reference:**

Traditionally the most common set up with a double junction reference electrode filled with the appropriate filling solution which is usually the ISAB or a dilution of the ISAB.

This set up is ideal in low concentration work where the outer filling solution of the reference can be diluted to minimise Interference or Ionic Strength problems. It is also advisable when several ISEs are being used simultaneously as this will be both economical and practical. Don't forget, if the reference fails it is fairly low cost to replace. Combination ISE's are more expensive and either a reference or sensor failure will mean a full electrode replacement.

Half cell ISE's are also easier to maintain or polish (crystalline membranes only) and therefore can have a longer lifetime.

Double junction reference electrodes are also slightly faster responding due to the liquid electrolyte, but do have the disadvantage of needing to be re filled and to be stored wet.

#### **Combination direction ISE with built in driTEK reference:**

This is the most practical way to perform Ion analysis in the field and is perfect for non skilled operators. The ISE and reference are combined in one body, which is stored dry and requires no or little maintenance.

The connector is a BNC which will fit all meters without having to worry about the configuration of the reference half cell connector.

Because the reference has a gel electrolyte it is both fairly concentrated and cannot be replenished. This means that working at low levels requires more care (particularly conditioning) and in these cases the Flow Plus or Half Cell options are preferable.

The reference electrolyte cannot be replenished and so the electrode lifetime is less than that of a Half Cell and a reference however the practicality of this set up outweighs any lifetime issues if you need to work in the field or require rapid analysis with unskilled operators.

**Flow Plus Combination ISE:**

This type of ISE combines the benefits of both of the above options. There is no need for a separate reference as it is a combined unit. It also has a liquid refillable reference which enables low concentration work to be performed but with the added advantage of being stored dry due to the ability to flush clean the electrode after use.

The sensor is also easy to maintain as the electrode can be taken apart and re assembled.

There are only two considerations which are; that you need a reference electrolyte for operation and that it is slightly more expensive than the conventional Combination ISE.

**For further information visit [www.edt.co.uk](http://www.edt.co.uk)**